Notiz / Note

Gas-Phase Generation and Characterization of Hydrogen Cyanide *N*-Methylide^[1]

Norman Goldberg, Muhammad Iraqi, and Helmut Schwarz*

Institut für Organische Chemie der Technischen Universität Berlin, Straße des 17. Juni 135, D-10623 Berlin, F.R.G.

Received July 23, 1993

Key Words: Ylides / Distonic ions / Neutralization-reionization mass spectrometry

The recently described (Maier et al.^[2]) title compound is not only viable in an argon matrix (10 K) but is also accessible in the gas phase. Electron impact ionization of $(CH_3)_3CCX_2NC$

The increasing importance and conspicuous properties of distonic ions^[3] – species in which the charge and radical centres are formally separated - as central intermediates in areas as diverse as chemistry in the gas phase or in frozen matrices are well documented in several review articles^[4]. An interesting sub-class are the a-distonic cations which formally correspond to the ionized forms of ylides^[5]. A typical example which highlights some of the unusual features of these molecules is $H_2\dot{C}-\dot{O}H_2$ (1). This radical cation has indeed remarkable properties in comparison with its conventional isomer, the methanol molecular ion CH₃OH^{+•} (2) (Scheme 1). In contrast to the neutral counterparts 5 and 6, ylide ion 1 is both thermochemically and kinetically more stable than 2^[6]. In charge stripping (CS) experiments^[7], 1 can be oxidized in the gas phase to its corresponding dication $H_2C-OH_2^{2+}$ (3), while CH_3OH^{+} upon gas-phase oxidation does not give rise to a detectable dication CH_3OH^{2+} (4)^[8]. Similarly, if $CH_2OH_2^{++}$ (1) is subjected to a neutralization-reionization (NR)^[9] experiment, in line with the theoretical predictions^[6d], no recovery signal for the *neutral* ylide H_2C-OH_2 (5) is observed^[10] in contrast to CH_3OH^{+} which, of course, can be neutralized to methanol. The failure to neutralize $H_2C-OH_2^+$ is twofold: (i) The carbon-oxygen binding energy of the neutral ylide 5 is quite small (<10 kcal/mol) and (ii) the large difference in equilibrium geometries of the charged and neutral species implies that vertically neutralized ylides will be formed with considerable excess

Scheme 1



(X = H, D) gives rise to $HCNCX_2^+$, which can be successfully neutralized in a beam experiment.

energy, thus facilitating rearrangement and/or fragmentation processes.

Here we report our results on the gas-phase generation and characterization of another important ylide, i.e. hydrogen cyanide Nmethylide HCNCH₂ (8) and its radical cation 7. The neutral species



Figure 1. Collisonal activation mass spectra (helium, 80% T) of (a) HCNCH₂⁺⁺, (b) HCNCD₂⁺⁺

8 is one of the members of the intriguing C_2H_3N family^[11], and in a recent exhaustive study^[2] its preparation in an argon matrix (10 K), spectroscopic indentification, and its relationship with other C_2H_3N isomers has been described.

Scheme 2



In the context of our earlier work on the identification of the potentially interstellar C₂H₃N isomer aminoacetylene^[12], we also reported on the collisional activation (CA)^[13] mass spectra of several $C_2H_3N^+$ isomers, including also HCNCH₂⁺ (7). This species is easily accessible by electron-impact-induced fragmentation of aliphatic isonitriles (Scheme 2), and a typical CA mass spectrum for this ion, generated from 2,2-dimethylpropyl isocyanide, is given in Figure 1a. In Figure 1b, the CA mass spectrum of the isotopomer HCNCD₂^{+•} is reproduced^[14].

The interpretation of Figure 1 is straightforward in that - in addition to the structure-unspecific losses of H_x (x = 1-3) – we observe three diagnostic features one would indeed expect for an HCNCX₂⁺ (X = H, D) connectivity: Formations of CX₂⁺, NCX₂⁺, and of the doubly charged ion $HCNCX_2^{2+}$. This signature is not only different from that of the other C₂H₃N^{+•} isomers, described



Figure 2. Neutralization-reionization mass spectra (xenon, 80% T/ oxygen 80% T) of (a) $HCNCH_2^+$, (b) $HCNCD_2^+$

earlier^[12], it is also in keeping with the connectivity of ionized hydrogen cyanide N-methylide 7.

Can 7 also be successfully neutralized? As evidenced by the spectra shown in Figure 2, the structure-indicative fragments are indeed retained, when HCNCX⁺ ions are subjected to a neutralizationreionization experiment. In addition, we observe a very intense survivor signal. These two observations clearly demonstrate that - in line with the theoretical expectations [2] - the neutral $HCNCX_2$ (X = H, D) is a stable molecule also in the gas phase. The relatively high abundances of the HCNCX₂ signals in the NR spectra are no surprise if one takes into account, that - in contrast to H_2O-OH_2 – no low-energy dissociation paths are available for hydrogen cyanide N-methylide (8). In particular, Maier's calculations^[2] indicate that the two carbon-nitrogen bonds of 8 are of nearly equal bond length, thus favouring the ylide form 8a in comparison to the N-methylide form 8b, which would be expected to fall apart more easily to HCN and CH₂.

$$\begin{array}{ccc} HC=N=CH_2 & HCN-CH_2 \\ 8a & 8b \end{array}$$

The generous financial support of our work by the Deutsche Forschungsgemeinschaft and the Fonds der Chemischen Industrie is appreciated. H. S. is grateful to the Alexander von Humboldt Foundation for a Max Planck Research Award which provides the financial background for the collaboration with his Israeli partner, Professor C. Lifshitz, from the Hebrew University of Jerusalem. We thank Professor G. Maier for providing us with a preprint of his recent work on C₂H₃N isomers in argon matrices.

- ^[2] G. Maier, C. Schmid, H. P. Reisenauer, E. Endlein, D. Becker, J. Eckwert, B. A. Hess, Jr., L. J. Schaad, Chem. Ber. 1993, 126,
- 2337-2352; preceding paper. ^[3] For a definition, see: L. Radom, W. J. Bouma, R. H. Nobes, B. F. Yates, *Pure Appl. Chem.* **1984**, *56*, 1831.
- F. Yates, Pure Appl. Chem. 1984, 36, 1831. [4] [4a] H. Schwarz, Nachr. Chem. Tech. Lab. 1983, 31, 451. [4b] H. Schwarz, Shitsuryo Burseki 1984, 32, 3. [4c] M. A. Baldwin, Specialist Periodical Report, Mass Spectrom. 1986, 8, 34. [4d] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [4e] G. Bouchoux, Mass Spectrom. Rev. 1988, 7, 123. [4d] K. M. Stirk, L. V. M. Sirik, L. V. Mass Spectrom. Rev. 1988, 7, 123. [40] M. Stirk, L. V. Mass Spectrom. Rev. 1988, 7, 123. [40] M. Stirk, L. V. Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 123. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom. Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. 1988, 7, 1203. [40] S. Hammerum, Mass Spectrom, Rev. K. M. Kiminkinen, M. I. Kenttämaa, Chem. Rev. 1992, 92, 1649.
- ^[5] D. G. Morris, Surv. Prog. Chem. 1983, 10, 189.
 ^[6] ^[6a] W. J. Bouma, R. H. Nobes, L. Radom, J. Am. Chem. Soc. 1982, 104, 2929. ^[6b] W. J. Bouma, J. K. MacLeod, L. Radom, J. Am. Chem. Soc. 1982, 104, 2930. ^[6c] J. L. Holmes, F. P. L. K. Tarlanni, P. C. Burgare, L. Am. Chem. Soc. 1982, 104, 2930. Lossing, J. K. Terlouw, P. C. Burgers, J. Am. Chem. Soc. 1982, 104, 2931. - ^[6d] B. F. Yates, W. J. Bouma, L. Radom, J. Am. Chem. Soc. 1987, 109, 2250.
- ^[7] W. Koch, F. Maquin, D. Stahl, H. Schwarz, Chimia 1985, 39, 376.
- ^[8] F. Maquin, D. Stahl, A. Sawaryn, P. v. R. Schleyer, W. Koch, G. Frenking, H. Schwarz, J. Chem. Soc., Chem. Commun. 1984, 504
- ^[9] Reviews: ^[9a] J. K. Terlouw, P. C. Burgers, B. L. M. van Baar, T. Weiske, H. Schwarz, *Chimia* **1986**, 40, 357. ^[9b] C. Wesdemiotis, F. W. McLafferty, *Chem. Rev.* **1987**, 87, 485. ^[9c] J. K. mous, r. w. McLanerty, Chem. Rev. 1987, 87, 485. – 64 J. K. Terlouw, H. Schwarz, Angew. Chem. 1987, 99, 829. Angew. Chem. Int. Ed. Engl. 1987, 26, 805. – ^[9d] H. Schwarz, Pure Appl. Chem. 1989, 61, 685. – ^[9e] J. L. Holmes, Mass Spectrom. Rev. 1989, 8, 513. – ^[91] J. K. Terlouw, Adv. Mass Spectrom. 1989, 11, 984. – ^[92] F. W. McLafferty, Science 1990, 247, 925. – ^[95] F. W. McLafferty, Int. J. Mass. Spectrom. Low Decomposed 1000 W. McLafferty, Int. J. Mass Spectrom. Ion Processes 1992, 118/119, 221
- ^[10] C. E. C. A. Hop, J. Bordas-Nagy, J. L. Holmes, J. K. Terlouw,
- Org. Mass Spectrom. 1988, 23, 155.
 ^[11] ^[114] H. Bock, B. Solouki, Angew. Chem. 1981, 93, 425; Angew. Chem. Int. Ed. Engl. 1981, 20, 427. ^[115] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 87, 355. ^[116] H. Dollarow Chem. 1981, 20, 427. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T.-K. Ha, M. T. Nguyen, J. Mol. Struct. (THEOREM) 1982, 77, 355. ^[116] T. Struct. (THEOREM) 1982, 77, 785. ^[116] T. Struct. (THEOREM) 1982, 787, 785. ^{[116} Bock, R. Dammel, S. Aygen, J. Am. Chem. Soc. 1983, 105, 7681.

^[1] Dedicated to Professor Hans Bock, Frankfurt, on the occasion of his 65th birthday.

- ^[11d] L. L. Lohr, Jr., M. Hanamana, K. Morokuma, J. Am. Chem. Soc. **1983**, 105, 5541. - ^[11e] S. Saeb, L. Farnell, N. V. Riggs, L. Radom, J. Am. Chem. Soc. **1984**, 106, 5047. - ^[110] M. Bogey, J.-L. Destombes, J.-M. Denis, J.-C. Guillemin, J. Mol. Spectrosc. **1986**, 115, 1. - ^[11g] H. Bock, R. Dammel, Angew. Chem. **1987**, 99, 518; Angew. Chem. Int. Ed. Engl. **1987**, 26, 504. - ^[11h] H. Bock, R. Dammel, Chem. Ber. **1987**, 120, 1971. -^[11i] J.-C. Guillemin, J.-M. Denis, M.-C. Lasne, J.-L. Ripoll, Tetrahedron **1988**, 44, 4447.

- [12] B. v. Baar, W. Koch, C. Lebrilla, J. K. Terlouw, T. Weiske, H. Schwarz, Angew. Chem. 1986, 98, 834; Angew. Chem. Int. Ed. Engl. 1986, 25, 827.
- ^[13] Review: K. Levsen, H. Schwarz, Angew. Chem. 1976, 88, 589; Angew. Chem. Int. Ed. Engl. 1976, 15, 509.
- ^[14] The mass-spectrometric experiments were performed using our modified ZAB-HF machine which is of *BEBE* configuration (*B* stands for magnetic and *E* for electric sector) (ref.^[15]). C₂H₃N⁺⁺ and C₂HD₂N⁺⁺ ions were generated by 70-eV ionization of (CH₃)₃CCH₂NC and (CH₃)₃CCD₂NC, respectively, employing the following ion-source conditions: ion-source temperature 200°C; trap current 100 μ A; repeller voltage 30 V; acceleration voltage 8 kV; mass resolution $m/\Delta m = 3000$. In the NRMS

experiment, a beam of B(1)E(1) mass-selected $C_2H_3N^{++}$, having 8 keV translational energy, was neutralized in the first cell of a differentially pumped tandem collision cell by colliding it with xenon (80% transmission, T). Unreacted ions were deflected away from the beam of neutral species by putting a voltage on a deflector electrode which is situated between the two collision cells. Subsequent reionization of the beam of neutral C_2H_3N is brought about in a second collision cell by collision with oxygen (80% T). The mass spectra of the resulting ions were recorded by scanning B(2). The *minimal* lifetime t (identical with the transit time from collision cell I to cell II) in the present experiments is >10 µs. In the CA experiments, helium was used as a collision partner (80% T). In order to improve the S/N ratio, signal averaging techniques were used by accumulating 25 - 50 scans and on-line processing the data with the AMD-Intectra data system.

system.
 ^[15] [^{15a]} R. Srinivas, D. Sülzle, W. Koch, C. H. DePuy, H. Schwarz, J. Am. Chem. Soc. 1991, 113, 5970. - [^{15b]} R. Srinivas, D. K. Bohme, D. Sülzle, H. Schwarz, J. Phys. Chem. 1991, 95, 9836. - [^{15c]} R. Srinivas, D. Sülzle, T. Weiske, H. Schwarz, Int. J. Mass Spectrom. Ion Processes 1991, 107, 369.

[239/93]